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(54) Title: PROCESS FOR THE PRODUCTION OF OLEFINS

(57) Abstract: A process for the production of an olefin from a hydrocarbon, which process comprises: partially combusting the hydrocarbon and an oxygen-containing gas in the presence of a catalyst, characterised in that the catalyst comprises platinum and at least one further metal, said further metal being a Group IIIA, Group IVA, VA or a transition metal; wherein said catalyst is: a) not a platinum catalyst consisting essentially of platinum modified with Sn, Cu or mixtures thereof, and b) not a platinum catalyst consisting essentially of platinum modified with Sb or a mixture of Sb and Sn.

### PROCESS FOR THE PRODUCTION OF OLEFINS

The present invention relates to a process for the production of olefins.

Olefins such as ethylene and propylene may be produced by a variety of processes, including the steam cracking of hydrocarbons or by the dehydrogenation of paraffinic feedstocks. More recently, olefins have been produced by a process known as auto-thermal cracking. In such a process, a hydrocarbon feed is mixed with an oxygen-containing gas and contacted with a catalyst. The hydrocarbon feed is partially combusted, and the heat produced is used to drive the dehydrogenation reaction.

An example of an auto-thermal cracking process is described in EP 0 332 289. The patent describes platinum group metals as being capable of supporting combustion beyond the fuel rich limit of flammability. Preferred catalysts are supported platinum catalysts such as platinum/gamma alumina spheres, and platinum/monoliths such as platinum/cordierite or mullite monoliths.

In WO 97/26987, such platinum catalysts are modified with Sn or Cu.

In WO 00/14035, platinum catalysts are modified with tin, copper or antimony. Specifically, Pt/Sn, Pt/Cu, Pt/Sb and Pt/Sn/Sb catalysts are disclosed.

We have now developed other catalysts for auto-thermal cracking.

Accordingly, the present invention provides a process for the production of an olefin from a hydrocarbon, which process comprises:

partially combusting the hydrocarbon and an oxygen-containing gas in the presence of a catalyst, characterised in that the catalyst comprises platinum and at least one further metal, said further metal being a Group IIIA, Group IVA, VA or a transition metal;

wherein said catalyst is a) not a platinum catalyst consisting essentially of platinum modified with Sn, Cu or mixtures thereof, and

b) not a platinum catalyst consisting essentially of platinum modified with Sb or a mixture of Sb and Sn.

5 It should be understood that unless otherwise specified, the term "further metal" covers all elements of Group IIIA, IVA, VA and the transition metal series of the Periodic Table.

For the avoidance of doubt, the platinum and at least one further metal may be present in the catalyst in any form, for example, in metallic form, or in the form of a metal  
10 compound, such as an oxide.

The partial combustion reaction is carried out by contacting a feed comprising the hydrocarbon and a molecular oxygen containing gas with the catalyst. Any suitable oxygen-containing gas may be employed; oxygen being preferred.

The preferred stoichiometric ratio of hydrocarbon to oxygen is 5 to 16,  
15 preferably, 5 to 13.5 times, more preferably, 6 to 10 times the stoichiometric ratio of hydrocarbon to oxygen required for complete combustion of the hydrocarbon to carbon dioxide and water.

Preferably, hydrogen is co-fed into the reaction. It is believed that in the presence of the catalyst, hydrogen combusts preferentially relative to the hydrocarbon, thereby  
20 increasing the olefin selectivity of the overall process.

Additional feed components such as nitrogen, carbon monoxide and steam may also be fed into the reaction.

Suitable Group IIIA metals include Al, Ga, In and Tl. Of these, Ga and In are preferred.

25 Suitable Group IVA metals include Ge, and Pb. Of these, Ge is preferred.

Suitable Group VA metals include Bi.

Suitable metals in the transition metal series are any metal from Group IB to VIIIB of the Periodic Table. In particular, transition metals selected from Groups IB, IIB, VIB, VIIB and VIIIB of the Periodic Table are preferred. Examples of such metals  
30 include Cr, Mo, W, Fe, Ru, Os, Co, Rh, Ir, Ni, Pd, Ag, Au, Zn, Cd and Hg. Preferred transition metals are Mo, Rh, Ru, Ir, Pd, and Zn.

In one embodiment of the present invention, the catalyst comprises only one metal

selected from Group IIIA, Group IVA, VA and the transition metal series. For example, the catalyst may comprise Pt and one of Ga, In, Ge and Bi.

In such an embodiment, platinum may nominally form between 0.01 and 5.0 wt %, preferably, between 0.01 and 2.0 wt %, and more preferably, between 0.05 and 1.0 wt % of the total weight of the catalyst. It should be noted, however, that not all the metal employed during the preparation of the catalyst necessarily becomes incorporated in the catalyst composition. Thus, the actual loading of metal may differ from the nominal loading. To ensure that the desired actual metal concentrations are achieved, the nominal metal concentrations may have to be varied accordingly.

10 The actual loading of platinum may be between 10 and up to 100% of the nominal loading. Preferably, the actual loadings are above 40%, more preferably, above 70 % (eg 90 to 99%) of the nominal values. In a preferred embodiment, platinum may actually form between 0.01 and 5.0 wt %, preferably, between 0.01 and 2.0 wt %, and more preferably, between 0.05 and 1.0 wt % of the total weight of the catalyst. The atomic  
15 ratio of platinum to the Group IIIA, IVA or transition metal may be 1 : 0.1 - 50.

Where a Group IIIA metal is employed, it is preferably gallium or indium. Preferred atomic ratios of Pt:Ga are 1:0.1 - 20.0, more preferably, 1: 0.2 - 15.0, even more preferably, 1:1.0 - 8.0, most preferably, 1:5.0 - 8.0, for example, 1:7.5. Preferred atomic ratios of Pt:In are 1:0.1 - 10.0, more preferably, 1: 0.2 - 7.0, and most preferably,  
20 1:0.8 - 3.0., for example, 1:1.3.

Where a Group IVA metal is employed, Ge is preferred. Preferred atomic ratios of Pt:Ge are 1: 0.1 - 20.0, more preferably, 1: 0.2 - 15.0 and most preferably, 1:0.5 - 12, for example, 1:10.8.

Where a Group VA metal is employed, Bi is preferred. Preferred atomic ratios of  
25 Pt:Bi are 1: 0.1 - 20.0, more preferably, 1: 0.2 - 15.0 and most preferably, 1:0.5 - 10.0.

Where a transition metal is employed, the metal is preferably Mo, Rh, Ru, Ir or Zn.

In another embodiment, the catalyst comprises at least two metals selected from Group IIIA, Group IVA, VA and the transition metal series. For example, the catalyst  
30 may comprise Pt, Pd and Sn or Pt, Pd and Cu. Platinum may nominally form between 0.01 and 5 wt %, preferably, between 0.01 and 2.0 wt %, and more preferably, 0.05 - 1.0 wt % of the total weight of the catalyst. As described above, the actual loading of metal

may differ from the nominal loading. Thus, the actual loading of platinum may be between 10 and up to 100% of the nominal value. Preferably, the actual loadings are above 40%, more preferably, above 70 % (eg 90 – 99%) of the nominal values. In a preferred embodiment, the actual loading of platinum is between 0.01 and 5 wt %, preferably, between 0.01 and 2.0 wt %, and more preferably, 0.05 - 1.0 wt % of the total weight of the catalyst.

Preferably, the catalyst comprises a) platinum, b) a further transition metal and c) a Group IIIA or IVA metal. The transition metal b) is preferably palladium. The metal c) is, preferably, a Group IVA metal, more preferably Sn or Ge. The atomic ratio of metal b) (eg palladium) to platinum may be 1: 0.1 - 10.0, preferably, 1: 0.5 - 8.0, and more preferably, 1: 1.0 - 5.0. The atomic ratio of metal b) (eg palladium) to metal c) (eg Sn or Cu) may be 1: 0.1 - 60, preferably, 1: 0.1 - 50, for example, 1: 0.1 - 12. Where metal b) is Pd and metal c) is Sn, the Pd:Sn atomic ratio is preferably 1: 0.1 - 60, more preferably, 1: 0.1 - 50, for example, 1: 0.1 - 12.0. Where metal b) is Pd, and metal c) Cu, the atomic Pd:Cu ratio is preferably 1: 0.1 - 3.0, more preferably, 1: 0.2 - 2.0, and most preferably, 1: 0.5 - 1.5.

The catalyst employed may be unsupported. For example, the catalyst may be in the form of a metal gauze. Preferably, however, the catalyst employed in the present process is a supported catalyst. The catalyst may be supported on any suitable support. Ceramic supports are generally preferred, although metal supports may also be employed.

Where ceramic supports are used, the composition of the ceramic support may be any oxide or combination of oxides that is stable at high temperatures of, for example, 600°C and 1200°C. The support material preferably has a low thermal expansion coefficient, and is resistant to phase separation at high temperatures.

Suitable ceramic supports include lithium aluminium silicate (LAS), alumina ( $\alpha$ - $\text{Al}_2\text{O}_3$ ), yttria stabilised zirconia, alumina titanate, niascon, and calcium zirconyl phosphate. The supports may be wash-coated, for example, with  $\gamma$ - $\text{Al}_2\text{O}_3$ .

The structure of the support material is important, as this may affect flow patterns through the catalyst. Such flow patterns may influence the transport of reactants and products to and from the catalyst surface, thereby affecting the catalyst's activity. Preferably, the substrate is a continuous multi-channel ceramic structure, such as a foam, a regular channelled monolith or a fibrous pad. The pores of foam monolith structures

tend to provide tortuous paths for reactants and products. Such supports may have 20 to 80, preferably, 30 to 50 pores per inch. Channel monoliths generally have straighter, channel-like pores. These pores are generally smaller, and there may be 80 or more pores per linear inch of catalyst. The support may be in the form of spheres or other granular shapes, or may be present as a thin layer or wash coat on another substrate.

The catalyst employed in the present invention may comprise further elements, such as alkali metals. Suitable alkali metals include lithium, sodium, potassium and caesium.

The catalyst employed in the present invention may be prepared by any method known in the art. For example, gel methods and wet-impregnation techniques may be employed. Typically, the support is impregnated with one or more solutions comprising the metals, dried and then calcined in air. The support may be impregnated in one or more steps. Preferably, multiple impregnation steps are employed. The support is preferably dried and calcined between each impregnation, and then subjected to a final calcination, preferably, in air. The calcined support may then be reduced, for example, by heat treatment in a hydrogen atmosphere.

In the case of a catalyst comprising platinum, palladium and tin, the support material is preferably impregnated in a platinum/palladium solution, and then in a solution comprising tin. Once impregnated, the catalyst is dried at eg 50 to 200°C, and calcined at eg 100 to 700°C e.g. 300 to 700°C, between each impregnation. The support is then subjected to a final calcination step at eg 400 to 1500°C in air. The catalyst may then be reduced in, for example, a hydrogen atmosphere. This reduction step may be carried out at temperatures of up to 1000°C, for example, 300 to 750 °C are employed.

The catalyst may be in the form of a fluidised or fixed bed. Preferably, a fixed bed catalyst is employed.

The partial combustion reaction may be suitably carried out at a catalyst exit temperature of between 600°C and 1200°C, preferably between 850°C and 1050°C and most preferably, between 900°C and 1000°C.

The reaction may be carried out at atmospheric or elevated pressure. Suitable pressures range from 0 to 2 bara, preferably 1.5 to 2 bara, for example 1.8 bara. Elevated pressures of for example, 2 to 50 bara, may also be suitable.

The hydrocarbon may comprise any suitable hydrocarbon. Preferably, gaseous

hydrocarbons are employed. Suitable gaseous hydrocarbons include ethane, propane, butane and mixtures thereof.

The hydrocarbon is passed over the catalyst at a gas hourly space velocity of greater than  $10,000 \text{ h}^{-1}$ , preferably above  $20,000 \text{ h}^{-1}$  and most preferably, greater than  
5  $100,000 \text{ h}^{-1}$ . It will be understood, however, that the optimum gas hourly space time velocity will depend upon the pressure and nature of the feed composition.

Advantageously, heat may also be supplied by pre-heating the hydrocarbon. The temperature to which the hydrocarbon, oxygen-containing gas and (optionally) hydrogen mixture may be preheated, however, is limited by the autoignition properties (eg  
10 temperature) of the feed.

Where the cracking reaction is carried out at elevated pressure, the reaction products may be quenched as they emerge from the reaction chamber to avoid further reactions taking place.

Any coke produced in the process of the present invention may be removed by  
15 mechanical means, or using one of the decoking methods described in EP 0 709 446.

These and other aspects of the present invention will now be described, by way of example, with reference to Figure 1 of the drawings, which is a schematic view of an apparatus suitable for carrying out an embodiment of the present invention.

Figure 1 depicts an apparatus 10 comprising a reactor 12 surrounded by a  
20 furnace 14. The reactor 12 may be formed of quartz or metal. Where a metal reactor 12 is employed, the inside of the reactor is lined with quartz (not shown). Typically, metal reactors are more susceptible to heat loss than quartz reactors.

The reactor 12 is coupled to an oxygen supply 16 and a hydrocarbon feed supply 18. The hydrocarbon feed comprises ethane, and hydrogen and nitrogen. A catalyst 20 is  
25 located within the reactor 12. The catalyst 20 is positioned between a pair of LAS heat shields 22, 24.

The furnace is set to minimise heat losses, and the reactants 16, 18 are introduced into the reactor via line 26. As the reactants contact the catalyst 20, some of the ethane in the hydrocarbon feed 18 combusts to produce water and carbon oxides. The hydrogen  
30 co-feed also combusts to produce water. Both these combustion reactions are exothermic, and the heat produced is used to drive the dehydrogenation of ethane to ethylene.

## Examples

### Example 1 - Preparation of Pd/Pt/Sn Catalyst

The catalyst was prepared by multiple impregnation of a lithium aluminium silicate support having a high purity alumina (HPA) wash-coat. The support was  
5 impregnated in 1) a platinum/palladium solution ( $(\text{NH}_3)_4\text{Pt}^{\text{II}}\text{Cl}_2$ ,  $(\text{NH}_3)_4\text{Pd}^{\text{II}}\text{Cl}_2$ ), and 2) an  $\text{SnCl}_2/\text{HCl}$  solution. Between each impregnation, the support was dried at  $120^\circ\text{C}$ , and calcined at  $450^\circ\text{C}$ . The catalyst was then calcined in air at  $600^\circ\text{C}$  for 6 hours, and then reduced in an atmosphere of hydrogen (1.0nl/min), and nitrogen (1.5 nl/min) for 1 hour (at  $700^\circ\text{C}$ ).

10 The catalyst was analysed and found to have 0.36 wt % Pt, 0.04 wt % Pd and 1.85 wt % Sn.

### Comparative Example A - 1 wt% Pt

A catalyst having a nominal loading of 1 wt % Pt was prepared by impregnating a lithium aluminium silicate support having an HPA wash-coat in a solution of  
15  $(\text{NH}_3)_4\text{Pt}^{\text{II}}\text{Cl}_2$ . The impregnated support was dried at  $120^\circ\text{C}$ , and calcined at  $450^\circ\text{C}$ . The catalyst was then calcined in air at  $1200^\circ\text{C}$  for 6 hours. The catalyst was analysed and found to have 0.86 wt % Pt.

### Comparative Example B - Pt/ Sn

A catalyst comprising Pt and Sn was prepared by impregnating a lithium  
20 aluminium silicate support having an HPA wash-coat in a solution of 1)  $(\text{NH}_3)_4\text{Pt}^{\text{II}}\text{Cl}_2$ , and 2)  $\text{SnCl}_2/\text{HCl}$ . The impregnated support was dried, calcined and reduced as described in connection with Example 1 above. The catalyst was analysed and found to have 0.48 wt % Pt, and 2.80 wt % Sn.

### Example 2

25 The catalysts of Example 1, Comparative Example A and Comparative Example B above were each tested as catalysts for the oxidative dehydrogenation of ethane. Each catalyst was mounted in the apparatus of Figure 1, and an oxidative dehydrogenation reaction was carried out under the conditions summarised in Table 1 below. For the tests, a metal reactor 12 was employed.

30



**Table 1**

	1 (Pd/Pt/Sn)	A (Pt only)	B (Pt/Sn)
GHSV @ stp/h	120863	119759	120268
ethane flow (g/min)	18.07	17.28	18.07
hydrogen flow (g/min)	1.18	1.21	1.18
oxygen flow (g/min)	9.39	9.68	9.39
nitrogen flow (g/min)	4.95	4.52	4.72

As shown in Table 2 below, the selectivity of the catalyst of Example 1 towards ethylene is greater than those of Comparative Examples A and B, respectively.

5

**Table 2**

	1 (Pd/Pt/Sn)	A (Pt only)	B (Pt/Sn)
ethane conversion (%)	74.70	75.69	73.03
selectivity (g ethylene per 100g ethane converted)	73.93	67.54	69.7

**Example 3 - Preparation of Pt/Ge catalyst**

The catalyst was prepared by sequential impregnation of 30ppi ceramic foam blocks (lithium aluminium silicate with alumina washcoat) with aqueous tetrammineplatinum(II) chloride and germanium tetrachloride in ethanol. Between 10 impregnations the blocks were dried in air at 120-140°C for ca. 30 minutes then calcined in air at 450°C for 30 minutes, cooled to room temperature and the impregnation procedure was repeated. Once all the impregnation solution had been absorbed onto the foam the blocks were calcined in air at 600°C for 6 hours.

15 Several of the calcined blocks were given a further air calcination at 1200°C for 6 hours before testing. Subsequent to air calcination at 600-1200°C the blocks could be reduced in a hydrogen (1.0 nl/min) and nitrogen (1.5 nl/min) atmosphere at 750°C for 1 hour prior to testing. The catalyst comprised 1.0 wt % Pt and 4.0 wt % Ge.

20

**Example 4**

- The catalyst of Example 3 was tested as a catalyst for the oxidative dehydrogenation of ethane using the apparatus of Figure 1. The reactor employed was formed of quartz. The performance of this catalyst was compared to that of Comparative Example A. The dehydrogenation conditions employed are summarised in Table 3 below. Table 4 shows the ethane conversion and selectivities achieved.

**Table 3**

	Pt -only Comp. Ex. A	1% Pt / 4% Ge
target (nominal) loadings Pt / Ge (wt%)	1.0 / ---	1.0 / 4.0
GHSV @ stp (l/h)	121073	120162
ethane flow (g/min)	18.94	18.94
hydrogen flow (g/min)	1.14	1.14
oxygen flow (g/min)	9.09	9.09
nitrogen flow (g/min)	5.00	4.68

10 **Table 4**

	Pt -only	1% Pt / 4% Ge
ethane conversion (%)	73.55	73.45
ethylene selectivity (g per 100g ethane converted)	71.44	73.20

**Example 5 - Preparation of a Pt/Ga catalyst**

- The catalyst was prepared by sequential impregnation of 30ppi ceramic foam block (99.5% alumina with alumina washcoat) with aqueous tetrammineplatinum(II) chloride and aqueous gallium nitrate. The foam block employed was 28mm diameter and 30mm in

depth. Between impregnations, the block was dried in air at 120-140°C for ca. 30 minutes, calcined in air at 450°C for 30 minutes, and then cooled to room temperature. Once all the impregnation solution had been absorbed onto the foam block, the block was calcined in air at 600°C for 6 hours.

- 5           The Pt and Ga solutions were prepared with sufficient salt to achieve final Pt/Ga loadings of 0.28 wt% platinum and 0.75 wt% gallium. The atomic platinum:gallium ratio was 1:7.5.

Prior to testing the catalysts were given an *in situ* reduction at 750°C under flowing hydrogen (ca. 1.0 nl/min) and nitrogen (ca. 1.5 nl/min) for 1 hour.

10    **Example 6**

The catalysts of Example 5, and Comparative Example A were tested as catalysts for the oxidative dehydrogenation of ethane. Each catalyst was mounted in the apparatus of Figure 1, and an oxidative dehydrogenation reaction was carried out under the conditions summarised in Table 5 below.

15

**Table 5**

	5 (Pt/Ga)	A (Pt only)
GHSV @ stp/h	121172	120130
ethane flow (g/min)	14.15	13.50
hydrogen flow (g/min)	12.73	13.17
oxygen flow (g/min)	6.37	6.58
nitrogen flow (g/min)	4.06	3.74

As shown in Table 6 below, the selectivity of the catalyst of Example 5 towards ethylene is greater than that of Comparative Example A.

20    **Table 6**

	5 (Pt/Ga)	A (Pt only)
ethane conversion (%)	77.71	79.64
selectivity (g ethylene per 100g	71.85	68.96

ethane converted)		
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### **Example 7 - Preparation of a Pt/In catalyst**

The catalyst was prepared by sequential impregnation of 30ppi ceramic foam blocks (99.5% alumina with alumina washcoat) with aqueous tetrammineplatinum(II) chloride and aqueous indium nitrate. The foam block used was 28mm in diameter and 30mm in depth. Between impregnations, the block was dried in air at 120-140°C for ca. 30 minutes, calcined in air at 450°C for 30 minutes, and then cooled to room temperature. Once all the impregnation solution had been absorbed onto the foam the blocks were calcined in air at 600°C for 6 hours.

The Pt and In solutions were prepared with sufficient salt to achieve final Pt/In loadings of 0.54 wt% platinum and 0.4 wt% indium. The catalyst had an atomic platinum:indium ratio of 1:1.26

Prior to testing the catalysts were given an *in situ* reduction at 750°C under flowing hydrogen (ca. 1.0 nl/min) and nitrogen (ca. 1.5 nl/min) for 1 hour.

### **Example 8**

In this Example, a Pt/In catalyst was prepared in the manner of Example 7 above, except that the catalyst was calcined in air at 1200°C for 6 hours, rather than at 600°C.

### **Example 9**

The catalysts of Example 7, 8 and Comparative Example A were tested as catalysts for the oxidative dehydrogenation of ethane. Each catalyst was mounted in the apparatus of Figure 1, and an oxidative dehydrogenation reaction was carried out under the conditions summarised in Table 7 below.

**Table 7**

	7 (Pt/In)	8 (Pt/In)	A (Pt only)
GHSV @ stp/h	119975	119908	120130
ethane flow (g/min)	13.50	13.50	13.50
hydrogen flow (g/min)	13.17	13.17	13.17
oxygen flow (g/min)	6.58	6.58	6.58
nitrogen flow (g/min)	3.69	3.67	3.74

As shown in Table 8 below, the catalysts of Examples 7 and 8 show higher selectivities towards ethylene than the catalyst of Comparative Example A.

**Table 8**

	7 (Pt/In)	8 (Pt/In)	A (Pt only)
ethane conversion (%)	78.86	77.57	79.64
selectivity (g ethylene per 100g ethane converted)	70.19	71.27	68.96

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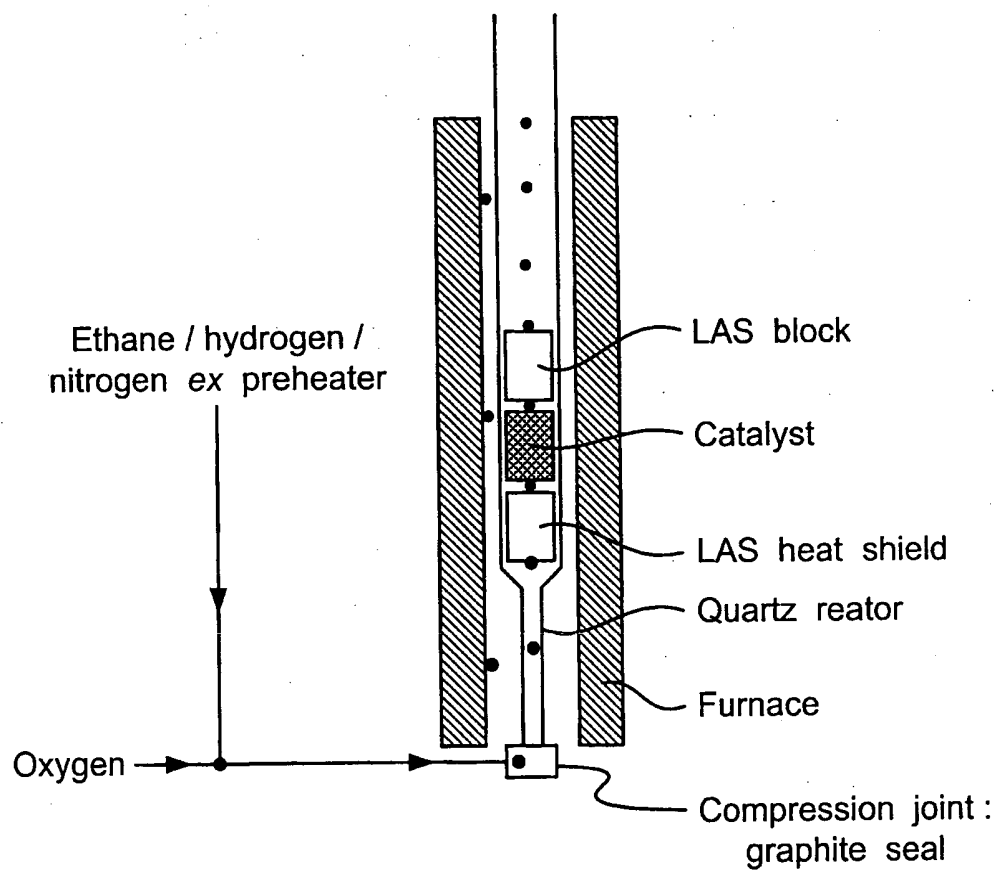
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## Claims

1. A process for the production of an olefin from a hydrocarbon, which process comprises:
  - partially combusting the hydrocarbon and an oxygen-containing gas in the presence of a catalyst, characterised in that the catalyst comprises platinum and at least one further metal, said further metal being a Group IIIA, Group IVA, VA or a transition metal;
  - wherein said catalyst is:
    - a) not a platinum catalyst consisting essentially of platinum modified with Sn, Cu or mixtures thereof, and
    - 10 b) not a platinum catalyst consisting essentially of platinum modified with Sb or a mixture of Sb and Sn.
2. A process as claimed in claim 1, wherein stoichiometric ratio of hydrocarbon to oxygen is 5 to 13.5 times, preferably, 6 to 10 times the stoichiometric ratio of hydrocarbon to oxygen required for complete combustion of the hydrocarbon to carbon dioxide and water.
- 15 3. A process as claimed in claim 1 or 2, wherein said partial combustion of said hydrocarbon and oxygen-containing gas is carried out in the presence of hydrogen.
4. A process as claimed in any preceding claim, wherein said catalyst comprises at least one further metal which is a Group IIIA metal.
- 20 5. A process as claimed in claim 4, wherein said Group IIIA metal is either Ga or In.
6. A process as claimed in any preceding claim, wherein said catalyst comprises at least one further metal which is a Group IVA metal.

7. A process as claimed in claim 6, wherein said Group IVA metal is Ge or Pb.
8. A process as claimed in any preceding claim, wherein said catalyst comprises at least one further metal which is a Group VA metal.
9. A process as claimed in claim 8, wherein said Group VA metal is Bi.
- 5 10. A process as claimed in any preceding claim, wherein said catalyst comprises at least one further metal which is a transition metal.
11. A process as claimed in claim 10, wherein said transition metal is selected from the group consisting of Cr, Mo, W, Fe, Ru, Os, Co, Rh, Ir, Ni, Pd, Ag, Au, Zn, Cd and Hg.
- 10 12. A process as claimed in any preceding claim, wherein said catalyst comprises only one said further metal.
13. A process as claimed in any one of claims 1 to 11, wherein said catalyst comprises at least two further metals selected from Group IIIA, Group IVA, VA and the transition metal series.
- 15 14. A process as claimed in claim 13, wherein the catalyst comprises one metal selected from Groups IIIA, IVA and VA, and one metal selected from the transition metal series.
15. A process as claimed in claim 14, wherein said metal selected from the transition metal series is Pt.
- 20 16. A process as claimed in claim 14 or 15, wherein said one metal selected from Groups IIIA, IVA and VA is Sn.
17. A process as claimed in claim 13, wherein said catalyst comprises two metals selected from the transition metal series.
18. A process as claimed in claim 17, wherein said catalyst comprises Pd, Pt and Cu.
- 25 19. A process as claimed in any preceding claim, wherein said catalyst further comprises alkali metal ions.
20. A process as claimed in any preceding claim, wherein said hydrocarbon comprises at least one hydrocarbon selected from the group consisting of ethane, propane and butane.

1/1



Key : • Location of thermocouples



# INTERNATIONAL SEARCH REPORT

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## A. CLASSIFICATION OF SUBJECT MATTER

IPC 7 C07C5/48

According to International Patent Classification (IPC) or to both national classification and IPC

## B. FIELDS SEARCHED

Minimum documentation searched (classification system followed by classification symbols)

IPC 7 C07C

Documentation searched other than minimum documentation to the extent that such documents are included in the fields searched

Electronic data base consulted during the international search (name of data base and, where practical, search terms used)

## C. DOCUMENTS CONSIDERED TO BE RELEVANT

Category *	Citation of document, with indication, where appropriate, of the relevant passages	Relevant to claim No.
X	WO 94 04632 A (BP CHEMICALS) 3 March 1994 (1994-03-03) claims	1,10-12, 20
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☐ Further documents are listed in the continuation of box C.

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